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Chemistry: The Central Science (13th Edition) Chapter 22 –

704 CHAPTER 22 Identify the product that balances the following nuclear reaction: $212\ 84\text{Po} \rightarrow 212\ 84\text{Pb} + 1\ 0\ 0\ \text{He}$. The total mass number and atomic number must be equal on both sides of the equation. $212\ 84\text{Po} \rightarrow 212\ 84\text{Pb} + 1\ 0\ 0\ \text{He}$ + mass number: 212 = 208 + 4 atomic number: 84 = 82 + 2. The nuclide has a mass number of 208 and an atomic number of 82, $208\ 82\text{Pb}$. 3.

CHAPTER 22 Nuclear Chemistry

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Chemistry (4th Edition) Burdge, Julia Publisher McGraw-Hill Publishing Company ISBN 978-0-07802-152-7

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Veritas/ Chemistry/ Chapter 4 Answer Key 3 Practice Problems 13. Use dimensional analysis and the given densities to make the following conversions. a. 14.8 g of boron (B) to cubic centimeters of boron. The density of boron is 2.34 g/ cm³. (14.8 ?)(1 3 2.34 ?)=6.32 3 b. 2.8 L of argon (Ar) to grams of argon.

Chapter 4 Problem Solving in Chemistry Answer Key

1) No, it is not a proper answer; you do not know whether the professor meant homework problem number 20 or 20 homework problems. Twenty is a quantity and does not make sense without a unit. 3) Middle Zeros: Zeros that appear between other nonzero digits are ALWAYS significant.

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