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PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0

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mol of N_2 are mixed with 12.0 mol of H_2 according to the following equation: $N_2(g) + 3H_2(g) \dots$

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compounds Reaction Stoichiometry - Involves
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for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$
How many moles of O_2 form if 3.0 mol of KClO_3 are totally consumed?
2. Given the following equation: $\text{H}_2(\text{g}) + \text{F}_2(\text{g}) \rightarrow 2\text{HF}(\text{g})$

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