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question. 1. In the reaction

$\text{CaCO}_3 + 2\text{HNO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 +$

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$\text{CO}_2 + \text{H}_2\text{O}$, how many
milliliters of 0.1015 M HNO_3
are required to react with
0.9857 grams of CaCO_3 ? (a)
0.1940

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$\text{Ca}(\text{NO}_3)_2 + \text{CO}_2 + \text{H}_2\text{O}$, how many milliliters of 0.1015 M HNO_3 are required to react

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Analysis Calculations
with 0.9857 grams of CaCO_3 ?

- (a) 0.1940 mL (b) 97.04 mL
(c) 194.1 mL 2.

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analysis) Answer key 1. Calculate the solubility product for SrF_2 given that a saturated solution is 8.6×10^{-4} M. $\text{SrF}_2 \rightleftharpoons \text{Sr}^{2+} + 2\text{F}^-$ $K_{sp} = [\text{Sr}^{2+}][\text{F}^-]^2$ $[\text{Sr}^{2+}] = [\text{SrF}_2] = 8.6 \times 10^{-4}$ $[\text{F}^-] = 2[\text{SrF}_2]$

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$2] = 2 \times 8.6 \times 10^{-4} = 17.2 \times 10^{-4} \text{ K sp}$
 $= 8.6 \times 10^{-4}$

$10^{-4} \times 17.2 \times 10^{-4} = 1.5 \times 10^{-10} \text{ 2.}$

Calcium fluoride, CaF_2 , is a sparingly ...

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